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162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write:
 $101.3 \text{ J} \cdot 0.18 \text{ L atm} \cdot 1 \text{ L atm}^{-1} = -18 \text{ J}$
6.17 An expansion implies an increase in volume, therefore w must be -325 J (see the defining equation for pressure-volume work.) If the system absorbs heat, q must be $+127 \text{ J}$. The change in energy (internal energy) is:

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Ch.6 - Thermochemistry Ch.6.1: The Nature of Energy Energy: An object's capacity to perform work or produce heat Potential Energy: Energy due to position or composition (chemical bonds). Kinetic Energy: Energy due to the motion of the object $\frac{1}{2} KE = mv^2$ Law of Conservation of Energy: Energy can neither be created nor destroyed,

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CHAPTER 6: THERMOCHEMISTRY To
convert the answer to joules, we write:
 $1013 \text{ J} \cdot 018 \text{ L atm} = 1 \text{ L atm} = - \cdot \times = \cdot w$
 $-18 \text{ J} \cdot 617$ An expansion implies an
increase in volume, therefore w must be
 -325 J (see the defining equation for
pressure-volume

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Thermochemistry is the study of the relationships between chemical reactions and energy changes. Kinetic Energy and Potential Energy Kinetic energy is the energy of motion: $E = \frac{1}{2}mv^2$ Potential energy is the energy an object possesses by virtue of its position.

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